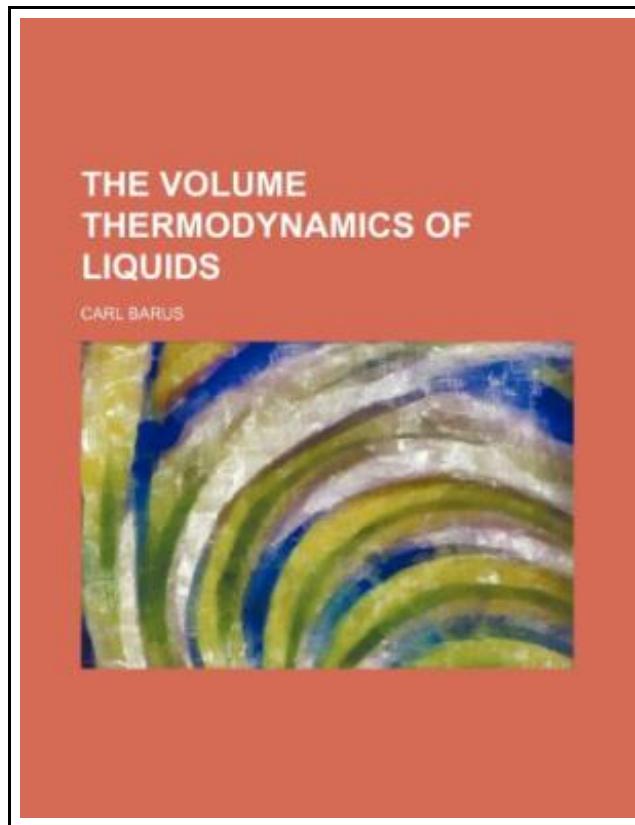


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1892 Excerpt: .at the same position from which Dupre and LeVy originally started. However, too much care can not be taken in keeping clearly in mind that pressures which, in relation to the usual laboratory facilities, are exceptionally large, may yet be mere vanishing increments when mapped out on the scale of the molecular pressures of liquids and solids. Some notion of the internal pressures encountered may be gained as follows: Studying the compression of organic substances throughout an interval of 310 and 600 atmospheres, I found that the isothermal compression v/V could be well represented by an equation $v/V = \ln(l+ap)/a$ where l and a are constants and p denotes pressure. Taking this as a basis, I sought the mean values of the constants involved in some fiftythree isothermals, and found $-v/V = \ln(1+9A)/9$ to apply within the limits of error of my work. To this extent, therefore, compressibility increases as the reciprocal of the first power of the pressure binomial $(l/9+a/p)$. The interpretation to be given to $l/9a$ is that of an initial or internal pressure. Computing its value for the above or organic liquids, I found the data of Table 30. Am. Jour., Sci. 3d ser., vol. 38, 1889, p. 407. 2 For a pressure interval not exceeding about 80 atmospheres for a group of isometrics, nor about 30 atmospheres for a single isometric. See Ranisay and Young, Phil. Mag.: 5th ser., vol. 33, 1887, p. 435. BAEus.j INTERNAL PRESSURE. 61 Table 30--Order of values for internal pressure,...

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